

Name: _____

Chem 9, Section: _____

Lab Partner: _____

Experiment Date: _____

Double Replacement Reactions

For each of the reactions performed: -- predict the reaction type (**precipitation**, **neutralization** or **gaseous**)
 -- record your observations
 -- predict the names and states of the products formed
 -- write the balanced equation, including all physical states.

1. Aqueous sodium chloride + aqueous silver nitrate

Reaction Type:	
Observations:	Product Names & States (if none, why not?):
Balanced Equation:	

2. Hydrochloric acid + aqueous sodium hydroxide

Reaction Type:	
Observations:	Product Names & States (if none, why not?):
Balanced Equation:	

3. Hydrochloric acid + solid sodium bicarbonate

Reaction Type:	
Observations:	Product Names & States (if none, why not?):
Balanced Equation:	

4. Aqueous iron (III) chloride + aqueous ammonium hydroxide

Reaction Type:	
Observations:	Product Names & States (if none, why not?):
Balanced Equation:	

5. Aqueous barium chloride + sulfuric acid

Reaction Type:	
Observations:	Product Names & States (if none, why not?):
Balanced Equation:	

6. Aqueous sodium phosphate + aqueous copper(II) sulfate

Reaction Type:	
Observations:	Product Names & States (if none, why not?):
Balanced Equation:	

7. Aqueous sodium carbonate + cobalt(II) nitrate

Reaction Type:	
Observations:	Product Names & States (if none, why not?):
Balanced Equation:	

8. Aqueous sodium chloride + aqueous potassium nitrate

Reaction Type:	
Observations:	Product Names & States (if none, why not?):
Balanced Equation:	

9. Aqueous nickel(II) nitrate + aqueous sodium hydroxide

Reaction Type:	
Observations:	Product Names & States (if none, why not?):
Balanced Equation:	

Describe each of the following compounds as ionic, covalent, or acid.
Use IUPAC nomenclature rules to write the chemical formula or the chemical name.

Name	Formula	Description
1. Aluminum oxide		
2. Calcium carbonate		
3. Tin(IV) sulfate		
4. Silver sulfite		
5. Nitric acid		
6. Carbon tetrachloride		
7. Hydrosulfuric acid		
8.	P_2O_5	
9.	NO_2	
10.	CaO_2	
11.	$PbCl_2$	
12.	$AlPO_4$	
13.	H_2SO_4	
14.	NBr_3	