

PRACTICE CHEMISTRY PLACEMENT EXAM with Weblinks for Self-Guided Tutorial

1. An atom with an electron configuration of $1s^2 2s^2 2p^3$ has how many valence electrons?

Khan Academy: [Electron Configurations](#); ChemTutor: [Atomic Structure](#)

- a. 2
- b. 3
- c. 4
- d. 5
- e. 7

2. The _____ sublevel of an atom has a total of five orbitals.

Khan Academy: [Quantum Numbers and Orbitals](#); ChemTutor: [Atomic Structure](#)

- a. s
- b. p
- c. d
- d. f
- e. g

3. You give a child a balloon, and he goes outside with it to play in the snow. Soon, he returns crying. What happened?

Khan Academy: [Ideal Gas Equation](#); ChemTutor: [Gases](#)

- a. The balloon expanded and burst.
- b. The balloon froze solid.
- c. The balloon shrank.
- d. The balloon dissolved.
- e. The child forgot Charles' Law.

4. What kind of bonding occurs in the compound potassium oxide?

Khan Academy: [Types of Chemical Bonds](#)

- a. ionic
- b. nonpolar covalent
- c. polar covalent (double bond)
- d. polar covalent (single bond)
- e. None of the above

5. A gas with a temperature of 21.0°C and a volume of 10.0 L is compressed to 5.00 L. What will be the new temperature?

Khan Academy: [Ideal Gas Equation](#); ChemTutor: [Gases](#)

- a. 10.5°C
- b. $420.^{\circ}\text{C}$
- c. 42.0°C
- d. -126°C
- e. 315°C

6. Which of the following sublevels does not exist as written?

Khan Academy: [Quantum Numbers and Orbitals](#)

- a. 3f
- b. 6f
- c. 2s
- d. 5d
- e. 8s

7. What is the molecular shape of PH_3 ?

Khan Academy: [Hybridization and Hybrid Orbitals](#)

- a. tetrahedral
- b. trigonal planar
- c. bent
- d. linear
- e. trigonal pyramidal

8. What is the percent by mass concentration of sodium bromide in a solution which contains 50.0 g of sodium bromide in 200.0 g of water?

ChemTutor: [Solutions](#)

- a. 40.0 %
- b. 20.0 %
- c. 25.0 %
- d. 33.3 %
- e. 50.0 %

9. How many milliliters of 6.00 M HCl solution would be required to prepare 2.00 L of 0.140 M HCl by dilution?

ChemTutor: [Solutions](#)

- a. 420 mL
- b. 168 mL
- c. 85.6 mL
- d. 46.7 mL
- e. 30.0 mL

10. What is the molar concentration of 2000. mL of aqueous solution containing 135 g of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$?

ChemTutor: [Solutions](#); Khan Academy: [Introduction to the Atom](#)

- a. 12.15 M
- b. 0.750 M
- c. 67.5 M
- d. 0.667 M
- e. 0.375 M

11. What is the formula of copper (II) sulfate pentahydrate?

ChemTutor: [Compounds](#)

- a. $\text{Cu}_2(\text{SO}_4)_2 \cdot 5 \text{H}_2\text{O}$
- b. $\text{Cu}_2(\text{SO}_4) \cdot 5 \text{H}_2\text{O}$
- c. $\text{CuSO}_4 \cdot 6 \text{H}_2\text{O}$
- d. $\text{CuSO}_4 \cdot 5 \text{H}_2\text{O}$
- e. None of the above

12. What is the electron configuration for the nitride ion?

Khan Academy: [Electron Configurations](#); ChemTutor: [Compounds](#)

- a. $1s^2 2s^2 2p^1$
- b. $1s^2 2s^2 2p^3$
- c. $1s^2 2s^2 2p^5$
- d. $1s^2 2s^2 2p^6$
- e. None of the above

13. A tank has a pressure of 30.0 atm at a temperature of 22.0°C. After heating, the temperature rises to 35.0°C. What is the new pressure?

Khan Academy: [Ideal Gas Equation](#); ChemTutor: [Gases](#)

- a. 54.3 atm
- b. 31.3 atm
- c. 28.7 atm
- d. 47.7 atm
- e. 30.6 atm

14. Which pair is immiscible?

ChemTutor: [Solutions](#); Khan Academy: [Solubility](#)

- a. ethanol and water
- b. water and octane, C_8H_{18}
- c. isopropyl alcohol and water
- d. acetic acid and water
- e. octane and oil

15. How many grams of sodium hydroxide are required to prepare 250.0 mL of a 6.00 M solution?

ChemTutor: [Solutions](#); Khan Academy: [Introduction to the Atom](#)

- a. 1.50 g
- b. 0.0375 g
- c. 0.600 g
- d. 3.75 g
- e. 60.0 g

16. 5.60 L of a gas at STP has a mass of 13.0 g. What is the molar mass of the gas?

Khan Academy: [Ideal Gas Equation](#); ChemTutor: [Gases](#)

- a. 33.2 g/mol
- b. 66.4 g/mol
- c. 26.0 g/mol
- d. 52.0 g/mol
- e. none of the above

17. What volume of 0.62 M sodium hydroxide is required to neutralize 20.00 mL of 0.391 nitric acid?

Word reaction with reactants *only*. (Students should predict products):

Sodium hydroxide + nitric acid

Khan Academy: [Balancing Chemical Equations](#); ChemTutor: [Reactions](#), [Stoichiometry](#)

- a. 23.6 mL
- b. 16.9 mL
- c. 9.03 mL
- d. 11.8 mL
- e. none of the above

18. How many moles are in 20.0 g of sodium carbonate?

Khan Academy: [Introduction to the Atom](#); ChemTutor: [Moles](#)

- a. 1.89 mol
- b. 212 mol
- c. 2.12×10^3 mol
- d. 0.189 mol
- e. 18.9 mol

19. The percent of nitrogen in magnesium nitride is

Khan Academy: [Molecular Composition](#); ChemTutor: [Moles](#), [Compounds](#)

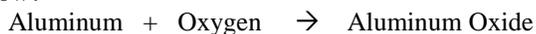
- a. 27.8%
- b. 36.6%
- c. 16.1%
- d. 72.2%
- e. 63.4%

20. What is the molar concentration of 250. mL of aqueous solution containing 48.8 g of glucose, $C_6H_{12}O_6$?

ChemTutor: [Solutions](#); Khan Academy: [Introduction to the Atom](#)

- a. 5.12 M
- b. 0.923 M
- c. 0.271 M
- d. 1.08 M
- e. 0.195 M

21. How many grams of aluminum metal will react with 0.0500 mole of oxygen gas according to the unbalanced equation given below?



Khan Academy: [Stoichiometry](#); ChemTutor: [Stoichiometry](#)

- a. 1.35 g
- b. 1.01 g
- c. 4.32 g
- d. 2.06 g
- e. 1.80 g

22. For the equation given, how many grams of methane will react with 125 g of oxygen?

Word reaction with reactants *only*. (Students should predict products):

Methane (CH₄) burns in oxygen

Khan Academy: [Stoichiometry](#); ChemTutor: [Stoichiometry](#)

- a. 39.1 g
- b. 19.5 g
- c. 15.6 g
- d. 31.3 g
- e. 62.5 g

For problems 23 - 24, Given the word reaction with reactants only (students should predict products):

phosphoric acid reacts with magnesium carbonate

23. From the balanced chemical equation the simplest whole number coefficient for the product magnesium phosphate is:

Khan Academy: [Balancing Chemical Equations](#); ChemTutor: [Reactions](#)

- a. 1
- b. 2
- c. 3
- d. 4
- e. none of the above

24. If 50.0 g of magnesium carbonate reacts completely with phosphoric acid, the grams of gas produced is

Khan Academy: [Balancing Chemical Equations](#); ChemTutor: [Reactions](#), [Stoichiometry](#)

- a. 52.2 g
 - b. 26.1 g
 - c. 13.1 g
 - d. 50.0 g
 - e. 55.0 g
-

25. How many molecules are in 5.8 g of acetone, C_3H_6O ?

Khan Academy: [Atomic Mass and Moles](#); ChemTutor: [Moles](#)

- a. 0.10 molecules
- b. 6.0×10^{22} molecules
- c. 3.5×10^{24} molecules
- d. 6.0×10^{23} molecules
- e. none of the above

26. This reaction is an example of which of the following types?

aluminum reacts with bromine to produce aluminum bromide

ChemTutor: [Reactions](#)

- a. combination
- b. single displacement
- c. decomposition
- d. gaseous
- e. precipitation

27. What is the simplest whole number coefficient for aluminum bromide in the above reaction (#26)?

Khan Academy: [Balancing Chemical Equations](#); ChemTutor: [Reactions](#)

- a. 1
- b. 2
- c. 3
- d. 4
- e. none of the above

28. How many moles of oxygen are required for the complete reaction of 45g of C_2H_4 when it is burned?

Khan Academy: [Balancing Chemical Equations](#), [Stoichiometry](#);

ChemTutor: [Reactions](#), [Stoichiometry](#)

- a. 1.3×10^2 mol
- b. 0.64 mol
- c. 112.4 mol
- d. 4.8 mol
- e. none of the above

29. If 14.0 g of C_2H_4 is burned and the actual yield of water is 7.84 g, the percent yield in the reaction is:

Khan Academy: [Balancing Chemical Equations](#), [Stoichiometry](#);

ChemTutor: [Reactions](#), [Stoichiometry](#)

- a. 0.56%
 - b. 43.6%
 - c. 87.1%
 - d. 56.0%
 - e. 82.0%
-

Answers:

1d 2c 3c 4a 5d 6a 7e 8b 9d 10e

11d 12d 13b 14b 15e 16d 17e 18d 19a 20d

21e 22d 23a 24b 25b 26a 27b 28d 29b